**Intro**

ACIDIC Hs WILL BE INDICATED ON EXAM IF UNCLEAR

STRENGTH OF ACIDS ≈ STABILITY OF CONJ. BASE

A) SAME PERIOD: MORE eNEGA ATOM W/ NEGA CHARGE
   SAME COLUMN: LARGER ATOM W/ NEGA CHARGE

B) RESONANCE STRUCTURES (MORE IS BETTER) → EXPLAIN EWG (EA NITRO)

C) INDICTION (EA - F)

D) ORBITALS W/ MORE ‘S CHARACTER’ (SP > SP^2 > SP^3)
   (BECAUSE S ORBITAL HOLDS E^- MORE TIGHTLY)

**Lewis Acids/Bases:** ePair ACCEPTORS/DONORS, OFTEN 1 PRODUCT + NO CONJ A/B

**Ex from Lecture:**

Identify the Lewis acid and the Lewis base in the reaction between BH₃ and THF.

BH₃ is the electron-pair acceptor (Lewis acid), and THF is the electron-pair donor (Lewis base).