Worksheet #22 Chem177 Awino

How to draw lewis structures:
1.) Draw each element
2.) The element with the highest bonding capacity will go in the middle if there are more than two elements.
3.) Draw the dots around each element
4.) Connect lone electrons between elements

Example:
1.) Draw the lewis structure for NaCl
2.) Draw the lewis structure for NH$_3$
3.) Draw the Lewis structure for MgBr$_2$

Lattice Energy
- What is Lattice energy?
- Is lattice energy more likely to be endothermic or exothermic?
- What is Lattice Energy dependent upon?

Electronegativity
- What is electronegativity?
- What is the trend on the periodic table for electronegativity?
- Change in electronegativity determines polarity
- Explain nonpolar covalent, polar covalent, and ionic in terms of electron sharing

Example:
1.) Label each as polar covalent, nonpolar covalent, and ionic
   - F$_2$, Br$_2$, NaCl, HCl, HF,